

Measurement of Activity of Water in Aqueous Poly(ethylene glycol) Solutions (Effect of Excess Volume on the Flory–Huggins χ -Parameter)

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Activities of water are reported for aqueous solutions of poly(ethylene glycol) (PEG) with number average molecular weights of 300, 400, 4000, and 6000 are measured at 35, 45, 55, and 65 °C and at various concentrations from 0 to 40 wt % of PEG using vapor-pressure osmometry. A cubic equation in terms of PEG concentrations is sufficient for correlation of activities. Using measured activities, the Flory–Huggins interaction parameters have been calculated. The effect of temperature, molecular weight of PEG and its concentration, and volume change upon mixing on interaction parameter have been considered. Ignoring the volume change of mixing can lead to an average error of 2.5% with a maximum error of 5.3%.

Introduction

Mixtures of poly(ethylene glycol) (PEG) + Dextrane (DX) + water form a two-phase liquid–liquid system. The upper liquid phase is rich in PEG, and the lower phase is rich in DX. Proteins, nucleic acids, enzymes, and other biomolecules have different solubilities in these phases and can be separated from each other (Albertsson, 1986; Zaslavsky, 1995). In thermodynamic modeling for such separation methods, activities are the essential variables. There are different methods to measure the activities of components in these systems, such as differential vapor pressure (Haynes et al., 1989), improved isopiestic methods (Ochs et al., 1990), vapor pressure osmometry, and membrane osmometry (Gaube et al., 1993).

In the present report, activities of water in different solutions of PEG having different molecular weights are measured by vapor pressure osmometry at various temperatures and compositions. The results are given by the Flory–Huggins (FH) equation (Flory, 1953)

$$\ln a_1 = \ln \varphi_1 + \left(1 - \frac{1}{r_2}\right)(1 - \varphi_1) + \chi_{12}(1 - \varphi_1)^2 \quad (1)$$

where a_1 and φ_1 are activity and volume fraction of solvent and r_2 is the number of segments of polymer defined as the molar volume of polymer divided by the molar volume of solvent. χ_{12} is the interaction parameter of the system.

According to Flory and Huggins (Flory, 1953), χ_{12} should be independent of concentration and molecular weight of polymer. However, in applying the FH equation to almost all polymer solutions, especially polar systems, it is necessary to consider χ_{12} dependent on concentration and molecular weight of the polymer, in addition to temperature, to fit the activity data of components of the solution (Kamide et al., 1985; Qian et al., 1991).

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The polymer solution systems of interest in this work are aqueous solutions of PEG in which both PEG and water are polar and there is a change of volume upon mixing. Therefore the effects of temperature, concentration, molecular weight of polymer, and volume change of mixing must be considered by a person who wants to calculate interaction parameter.

Experimental Procedure

Samples of poly(ethylene glycol)s (PEG 300, PEG 400, PEG 4000, and PEG 6000) and urea, the calibration substance, were obtained from Merck. Double-distilled water was used in making up the aqueous polymer solutions. Solutions were prepared by mass, using an analytical balance with ± 0.1 mg accuracy (Shimadzu, model AEU210). A vapor pressure osmometer (VPO) manufactured by Knauer (Germany) was used for measuring the activities of water in the aqueous PEG solutions. This apparatus consists of a chamber filled with a saturated vapor of solvent. Two thermistors are provided. A droplet of solvent is placed on one of the thermistors, and a droplet of polymer solution is placed on the other. Since the vapor pressure of solvent and polymer solution differ from each other, the solvent vapor condenses onto the polymer solution droplet and causes a change in its temperature. Temperature differences between the two thermistors are represented as voltage difference measured by a digital voltmeter. The voltage difference has the following relation with the activity of solvent (Brown, 1967)

$$\ln a_1 = - \frac{\Delta V M_1}{1000 I} \quad (2)$$

where a_1 is the activity of solvent, ΔV is the voltage difference, M_1 is the molecular weight of solvent, and I is the calibration constant. The calibration constant has various values at different temperatures that must be

Table 1. Measured Number Average Molecular Weights

	PEG 300	PEG 400	PEG 4000	PEG 6000
M_n	296	430	4237	5989

Table 2. Experimental Results of Activity of Water in Aqueous Solutions of PEG 300, PEG 400, PEG 4000, and PEG 6000 at Various Temperatures and Concentrations

w_1	$t/^\circ\text{C}$			
	35	45	55	65
	PEG 300			
0.6074	0.9410	0.9458	0.9499	0.9533
0.7013	0.9643	0.9670	0.9690	0.9707
0.8002	0.9815	0.9821	0.9829	0.9837
0.8920	0.9917	0.9915	0.9919	0.9924
	PEG 400			
0.6850	0.9691	0.9707	0.9720	0.9710
0.7549	0.9802	0.9820	0.9828	0.9828
0.8208	0.9878	0.9888	0.9890	0.9889
0.9377	0.9968	0.9972	0.9970	0.9971
0.9502	0.9974	0.9978	0.9976	0.9977
0.9695	0.9983	0.9987	0.9986	0.9987
0.9790	0.9991	0.9991	0.9990	0.9993
0.9880	0.9996	0.9996	0.9995	0.9997
	PEG 4000			
0.6051	0.9726	0.9766	0.9829	0.9846
0.7060	0.9878	0.9897	0.9920	0.9935
0.8006	0.9953	0.9960	0.9968	0.9975
0.9510	0.9996	0.9997	0.9995	0.9997
	PEG 6000			
0.6448		0.9832	0.9872	
0.7069		0.9900	0.9925	0.9939
0.8026	0.9961	0.9960	0.9973	0.9977
0.9018	0.9992	0.9986	0.9995	0.9995
0.9207	0.9994	0.9989	0.9997	0.9996
0.9399	0.9996	0.9993	0.9998	0.9997
0.9601	0.9998	0.9998	0.9999	0.9998
0.9797	~1	~1	~1	~1

determined. The calibration material should be soluble in the solvent. Urea and NaCl are suggested by the manufacturer as appropriate calibration materials for aqueous solutions. When a droplet of solvent and a droplet of solution are placed on each of the thermistors, ΔV increases and reaches a maximum at the quasi steady state. The required time for reaching this condition depends on the type of solution and its temperature. However, usually 3–5 min is adequate. In this manner, one is able to measure the activities of solvents in aqueous polymer solutions with the molecular weight of polymer less than 10 000 and its concentration less than 40%. Each measurement is repeated five to eight times, and then the average voltage difference is used in eq 2. The accuracy of the ΔV measurements is estimated to be better than $\pm 1\%$.

Results and Calculations

Initially, number average molecular weights of PEG samples were measured by VPO. Results are reported in Table 1. Then, the solvent activities were measured by VPO for aqueous solutions of PEG300, PEG400, PEG4000, and PEG6000. Experimental results are reported in Table 2. The experimental activity data are well fitted to a third-order polynomial

$$a_1 = aw_1^3 + bw_1^2 + cw_1 + d \quad (3)$$

with respect to water mass fraction, w_1 . The coefficients of eq 3 along with its absolute percentage deviation are reported in Table 3. The low ARD% values reported in Table 3 indicate the high accuracy of eq 3 for activity correlation.

Table 3. Coefficients of Eq 3 for Various PEG Solutions and at Various Temperatures

$t/^\circ\text{C}$	a	b	c	d	ARD% ^a
	PEG 300				
35	0.3678	-1.1814	1.3223	0.4912	0.01
45	0.4471	-1.3269	1.3874	0.4925	0.01
55	0.3613	-1.0878	1.1621	0.5644	0.00
65	0.2692	-0.8426	0.9423	0.6321	0.00
	PEG 400				
35	0.3503	-1.1032	1.2019	0.5508	0.01
45	0.4591	-1.3845	1.4371	0.4884	0.01
55	0.5311	-1.5421	1.5434	0.4678	0.01
65	0.6469	-1.8548	1.8243	0.3839	0.02
	PEG 4000				
35	0.4149	-1.2333	1.2311	0.5873	0.00
45	0.3683	-1.0879	1.0796	0.6401	0.01
55	0.2490	-0.7373	0.7359	0.7524	0.01
65	0.3055	-0.8712	0.8352	0.7305	0.01
	PEG 6000				
35	0.3992	-1.1921	1.1920	0.6009	0.00
45	-0.4397	1.1361	-0.9518	1.2554	0.01
55	0.2723	-0.8177	0.8216	0.7239	0.01
65	0.4086	-1.1696	1.1202	0.6408	0.00

^a ARD% = $100 \times (\sum_{i=1}^n |a_1^{\text{cal}} - a_1^{\text{exp}}| / a_1^{\text{exp}}) / n$, where n is the number of experiments.

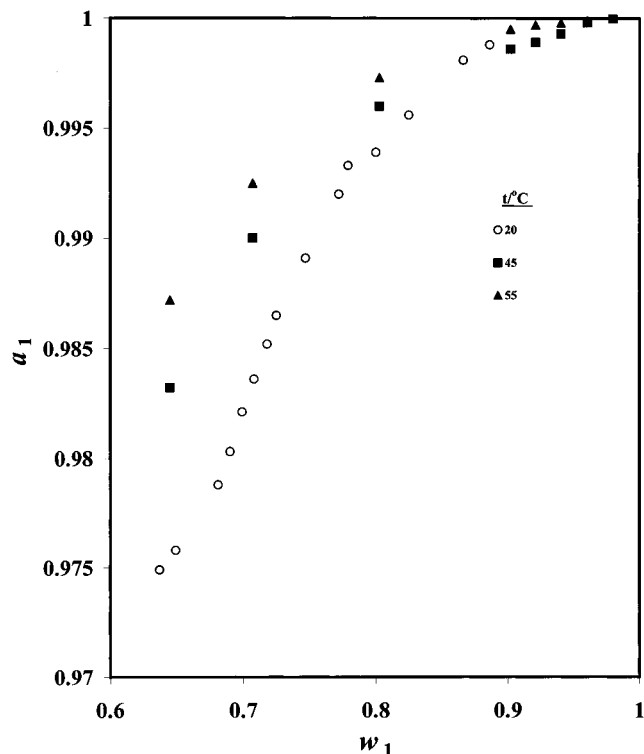


Figure 1. Activity of water in aqueous solution of PEG 6000. The empty circles are measured by Grossman et al. (1995), and the others are the present measurements.

Figures 1 and 2 shows activities measured in this work and also activities reported before (Grossman et al., 1995; Lin et al., 1996) for PEG 6000 and PEG 4000. The consistency between the results are very good.

The system of PEG + water has a volume change upon mixing. Using the density data (Eliassi et al., 1998), the excess volume

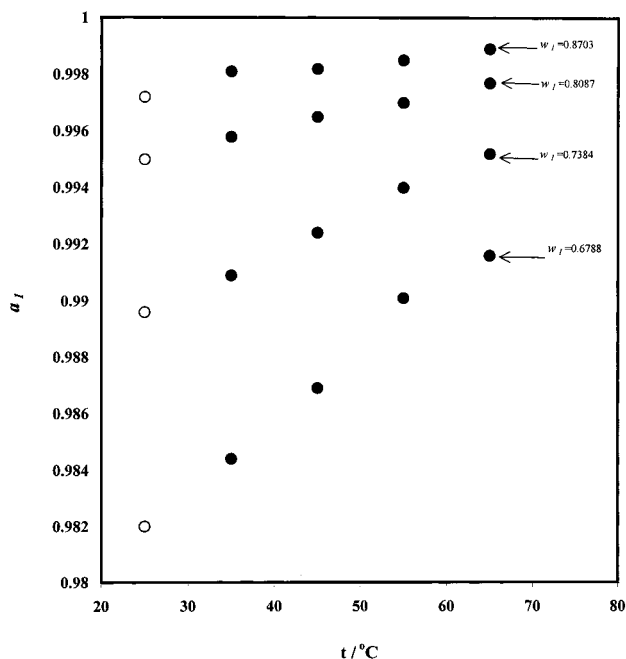
$$V^E = \frac{1}{\rho} - \left(\frac{W_1}{\rho_1} + \frac{W_2}{\rho_2} \right) \quad (4)$$

of PEG400 + water is calculated at various temperatures

Table 4. Flory–Huggins Interaction Parameters of PEG Aqueous Solutions at Various Temperatures and Mass Fractions of Polymer (w_2)

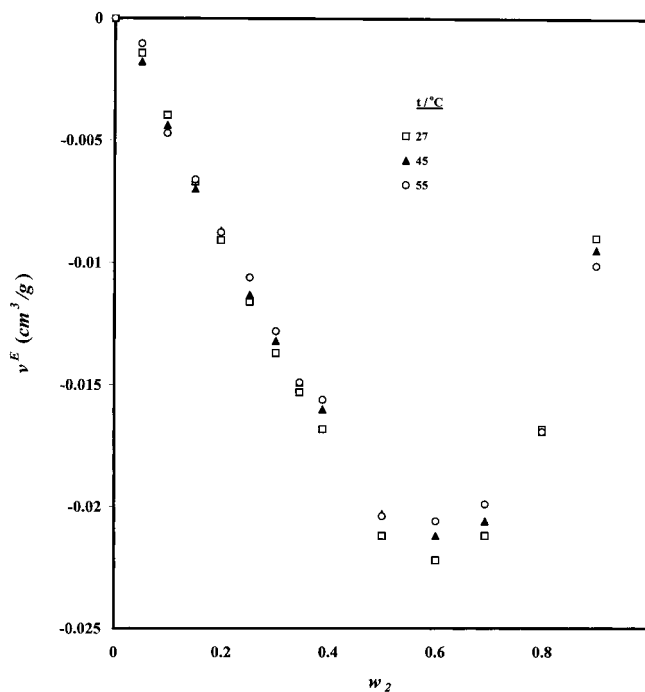
PEG 400									
		w_2							
$t/^\circ\text{C}$		0.0120	0.0210	0.0305	0.0498	0.0623	0.1792	0.2451	0.3156
45	$\tilde{\chi}_{12}$	1.8102	0.6400	0.6204	0.5442	0.5273	0.4376	0.4427	0.4321
	ARD% ^a	(5.28)	(0.31)	(1.56)	(0.68)	(0.38)	(2.42)	(2.91)	(3.59)
55	$\tilde{\chi}_{12}$	0.8801	0.3467	0.4765	0.4363	0.4583	0.4465	0.4608	0.4500
	ARD%	(3.17)	(2.83)	(0.17)	(1.06)	(0.75)	(2.20)	(2.47)	(3.15)
PEG 4000									
		w_2							
$t/^\circ\text{C}$		0.0490	0.1994	0.2940	0.3949				
45	$\tilde{\chi}_{12}$	0.4697	0.4653	0.4769	0.4923				
	ARD%	(1.10)	(2.62)	(3.21)	(3.77)				
55	$\tilde{\chi}_{12}$	0.3592	0.4925	0.5124	0.5470				
	ARD%	(3.72)	(0.95)	(2.33)	(2.19)				
PEG 6000									
		w_2							
$t/^\circ\text{C}$		0.0203	0.0399	0.0601	0.0793	0.0982	0.1973	0.2931	0.3552
45	$\tilde{\chi}_{12}$	0.3791	0.4443	0.3238	0.3386	0.3751	0.4496	0.4721	0.4865
	ARD%	(2.48)	(1.33)	(5.55)	(5.47)	(4.45)	(3.46)	(3.55)	(3.20)
55	$\tilde{\chi}_{12}$	0.3855	0.4475	0.4475	0.4901	0.5037	0.4962	0.4952	0.5095
	ARD%	(0.95)	(0.68)	(0.42)	(0.43)	(0.75)	(2.09)	(2.83)	(2.84)

$$^a \text{ARD\%} = 100 \times (\sum_{i=1}^n |\chi_{12} - \tilde{\chi}_{12}| / \tilde{\chi}_{12}) / n.$$

**Figure 2.** Activity of water in aqueous solutions of PEG 4000. The solid circles are from the present measurements. The empty circles are measured by Lin et al. (1996).

and concentrations. In eq 4, v^E is excess volume and ρ , ρ_1 , ρ_2 are densities of PEG solution, pure water, and pure PEG, respectively. Also, w_1 and w_2 are weight fractions of water and PEG in solution, respectively. Results of excess volume calculations are reported in Figure 3. According to this figure, at about 60 wt % of PEG concentration the system has the most negative excess volume at all temperatures.

Considering the above results it is clear that the volume change upon mixing is appreciable. Therefore, volume fractions of PEG are calculated in two ways, once using the partial volumes of water and PEG (Eliassi et al., 1998) and second using pure volumes instead of partial volumes. Also, the interaction parameter of the FH equation is

**Figure 3.** Excess volumes of the aqueous solutions of PEG 400.

calculated in two ways, namely, using the measured activities and volume fractions and ideal volume fractions, which are shown by $\tilde{\chi}_{12}$ and χ_{12} , respectively. In both cases, r_2 is calculated by dividing molar volume of polymer to molar volume of solvent (Eliassi et al., 1998). The values of these quantities accompanying the absolute percentage deviation (ARD%) are reported for PEG400, PEG4000, and PEG6000 in Table 4.

Conclusion

According to the measurements reported here water activity increases with increasing temperature, and the slope of the water activity versus water mass fraction

decreases with increase in PEG molecular weight. Also, variation of water activity with temperature in concentrated solutions is shown to be less than those of dilute solutions. It is also demonstrated that ignoring the volume change of mixing of PEG aqueous solutions can lead to an average error of 2.5% and a maximum error of 5.3% in calculating the Flory–Huggins interaction parameters. For the systems studied in this report it is also shown that the Flory–Huggins interaction parameters are dependent on the concentration and molecular weight of polymers considered.

Acknowledgment

The authors thank Dr. Mehdi Nekoumanesh and Ms. Roghieh Jamjah of the Polymer Research Center of Iran for their assistance and for providing access to their research facilities.

Literature Cited

- Albertsson, P. A. *Partition of Cell Particles and Macromolecules*, 3rd ed.; John Wiley and Sons: New York, 1986.
- Brown, W.; Solution Properties of Lignin. Thermodynamic Properties and Molecular Weight Determinations; *J. Appl. Polym. Sci.* **1967**, *11*, 2381–2396.
- Eliassi, A.; Modarress, H.; Mansoori, G. A.; Densities of Poly(ethylene glycol) + Water Mixtures in the 298.15–328.15 K Temperature Range. *J. Chem. Eng. Data* **1998**, *43*, 719–727.
- Flory, P. J. *Principles of Polymer Chemistry*; Cornell University Press: New York, 1953.
- Gaube, J.; Pfennig, A.; Stumpf, M.; Vapor–Liquid Equilibrium in Binary and Ternary Aqueous Solutions of Poly(ethylene glycol) and Dextran. *J. Chem. Eng. Data* **1993**, *38*, 163–166.
- Grossmann, C.; Tintinger, R.; Zhu, J.; Maurer, G. Aqueous Two-Phase Systems of Poly(ethylene glycol) and Dextran—Experimental Results and Modeling of Thermodynamic Properties. *Fluid Phase Equilib.* **1995**, *106*, 111–138.
- Haynes, R. A.; Beynon, R. S.; King, Blanch, H. W.; Prausnitz, J. M.; Thermodynamic Properties of Aqueous Polymer Solutions: Poly(ethylene glycol)/Dextran. *J. Phys. Chem.* **1989**, *93*, 5612–5617.
- Kamide, K.; Matsuda, S.; Satto, M.; Evaluation of Concentration of χ -Parameter, Flory Temperature and Entropy Parameter for Polymer–Solvent System from Their Critical Solution Temperature and Concentration Data. *Polym. J.* **1985**, *17*, 1013–1027.
- Lin, D. Q.; Zhu, Z. Q.; Mei, L. H.; Yang, L. R. Isopiestic Determination of the Water Activities of Poly(ethylene glycol) + Salt + Water Systems at 25 °C. *J. Chem. Eng. Data* **1996**, *41*, 1040–1042.
- Ochs, L. R.; Kabiri-Badr, M.; Cabezas, H., Jr. An Improved Isopiestic Method to Determine Activities in Multicomponent Mixtures. *AIChE J.* **1990**, *36*, 1908–1912.
- Qian, C.; Mumby, S. J.; Eichinger, B. E.; Existence of Two Critical Concentrations in Binary Phase Diagrams. *J. Polym. Sci., Part B* **1991**, *29*, 635–637.
- Zaslavsky, B. Y.; *Aqueous Two-Phase Partitioning*; Marcel Dekker Inc.: New York, 1995.

Received for review July 14, 1998. Accepted October 5, 1998. Financial support for this project was provided in part by the Iran Ministry of Culture and Higher Education.

JE980162Z